

## Lab 3. Hydrated Compounds

### Purpose:

To use hydrated compounds to practice calculating how much water molecules are attached to other molecules and compounds.

### Procedure:

1. Clean the crucible with water, dry it with a paper towel, and heat it for 2-3 minutes using the Bunsen burner.
2. Let the crucible cool for 3-5 minutes. Use the crucible tongs to carry the crucible to a balance and mass the empty crucible.
3. Mass out 3g of the assigned hydrated compound into the crucible.
4. Heat the crucible and its contents for about 10 minutes.
5. Let the crucible cool and re-mass it.
6. Heat the crucible and its contents for 2 more minutes.
7. Let the crucible cool and re-mass it.
8. If the last mass does not agree within 0.02g with the last mass reading, reheat the crucible and re-mass it until the last two measurements are within 0.02g of each other.
9. After final massing, place the crystals in the waste container provided. Clean the crucible, and clean the lab area.
10. Determine the percentage of the water in the assigned hydrated compound.

### Data:

Mass of crucible: 13.08g

Mass of hydrated compound: 3g  $\text{MgSO}_4 \cdot \text{H}_2\text{O}$

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