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Data:

	Trial 1	Trial 2	Trial 3	Trial 4	Trial 5	Trial 6
Amount of vinegar	10mL	10mL	10mL	10mL	10mL	10mL
Initial NaOH	1mL	9.9mL	18.0mL	26.2mL	34.4mL	36.65mL
Final NaOH	9.9mL	18.0mL	26.2mL	34.4mL	42.65mL	44.6mL

Observations:

First trial turned dark pink, too much NaOH. Accidentally let out a stream of NaOH near the endpoint. Trial 2 was a light pink, better than trial 1. Trial 3 was terrible. I accidentally added a stream of NaOH. Trial 4 was clear. half a drop turned it pink. Trial 5 was also a drop over from being clear and pink. Trial 6, you can barely tell that it is pink. If you look really close, you can see a very faint pink. It is really close to perfect. With some distilled water.

Analysis: $CH_3COOH+NaOH->CH_3COONa+H_2O$

Trial 1:

9.90mL-1.00mL=8.90mL NaOH used

1.00M*8.90mL/1000=0.0089 mol NaOH

0.0089 mol/10mL/1000=0.890M Acetic Acid

Trial 2:

18.0mL-9.9mL=8.1mL

(8.1/1000)/(10mL/1000)=0.81M

Trial 3:

26.2mL-18.0mL=8.2mL

(8.2mL/1000)/(10mL/1000)=0.82M

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