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7. Calculate the percent deviation between the trials.

$$\% \text{ deviation} = (\text{highest molarity} - \text{lowest molarity}) / \text{average molarity} * 100$$

8. If the % deviation is greater than 5%, repeat the titration process.

Data:

	Trial 1	Trial 2	Trial 3	Trial 4	Trial 5
KHP used	2.00g	2.00g	2.00g	2.00g	2.00g
Initial NaOH	0.20mL	10.00mL	19.70mL	29.4mL	39.15mL
Final NaOH	10.00mL	19.70mL	29.4mL	39.15mL	48.85mL

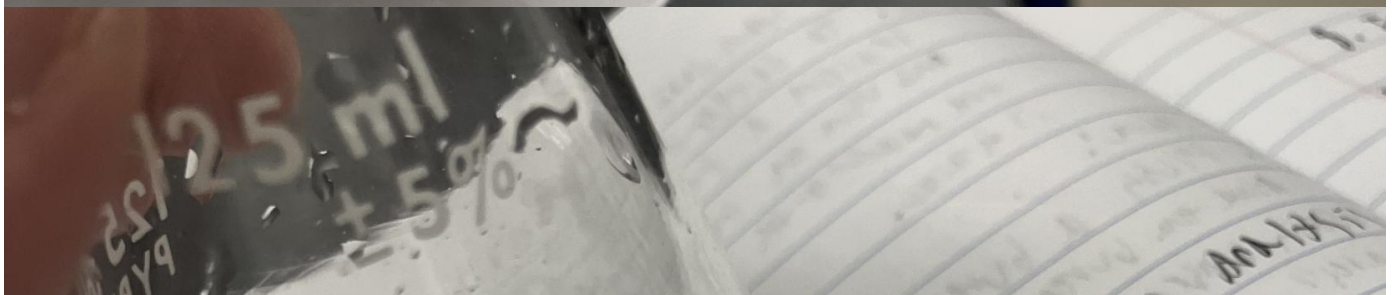
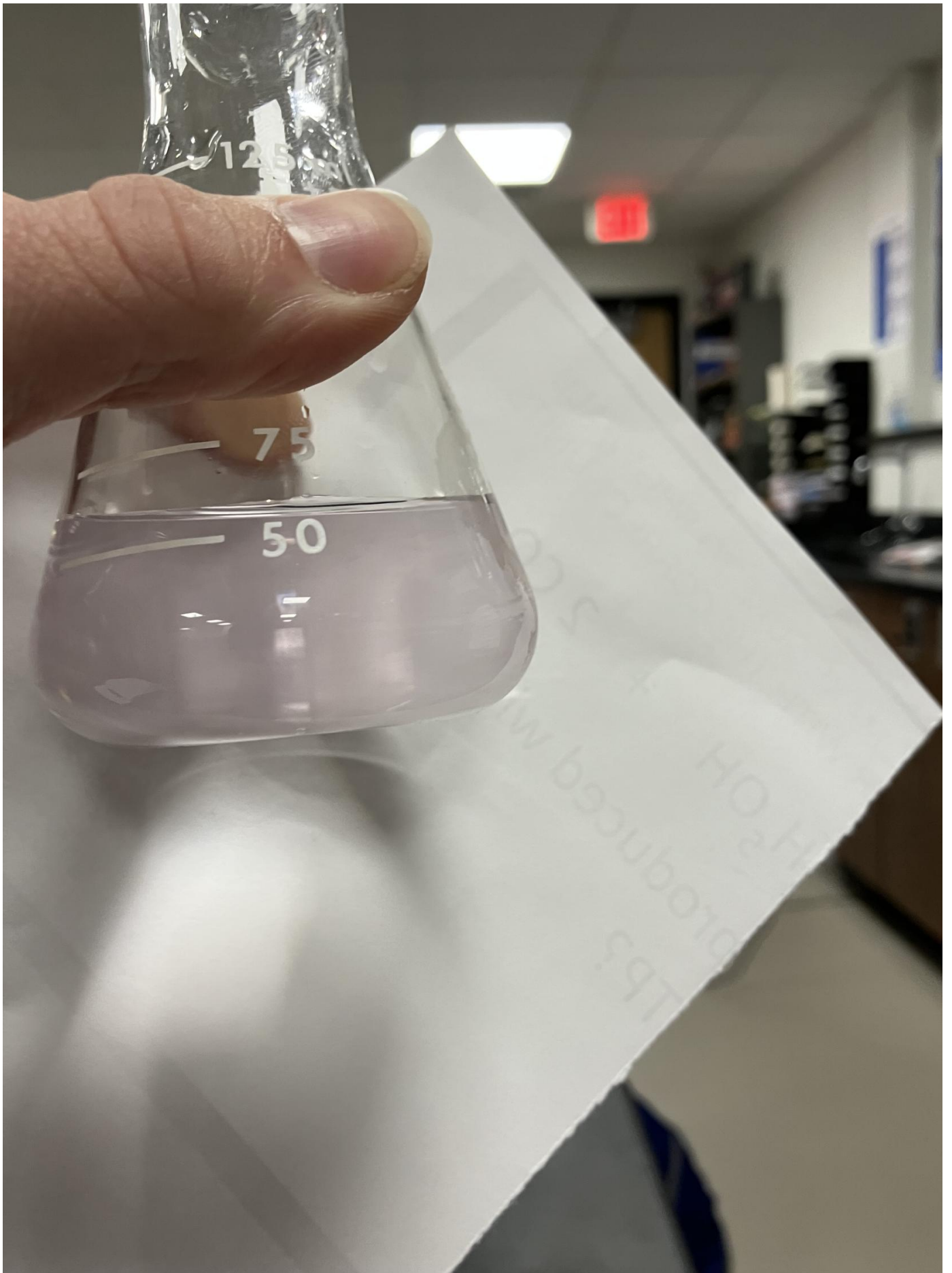
Observations:

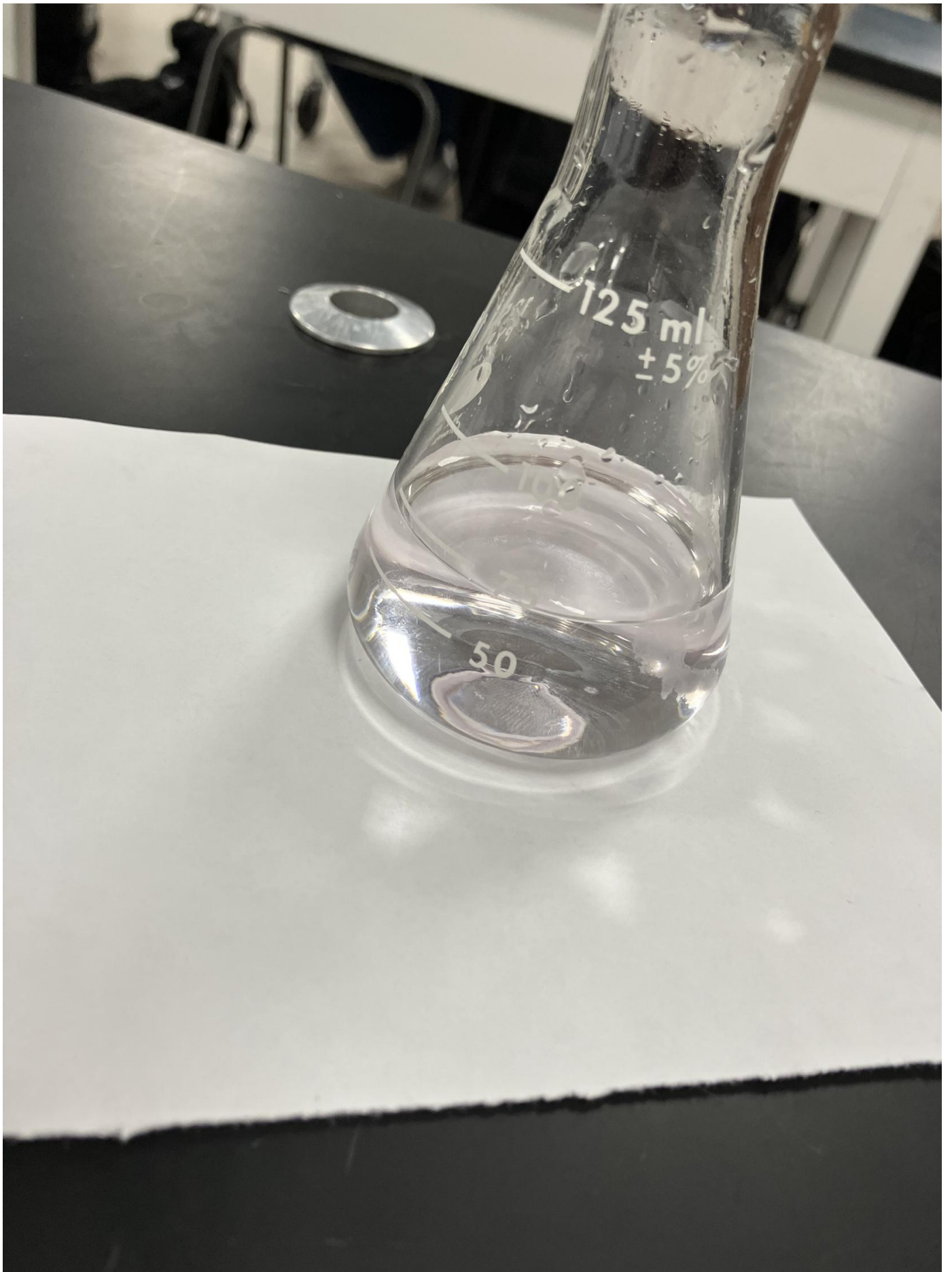
First trial turned a dark pink after about 10mL of NaOH. Way too dark. About a drop too much of NaOH. Trial 2 turned a dark pink, not as dark as trial 1. Trial 3 was similar to trial 2, same color and same amount of NaOH used. Magic number seems to be 9.7mL when it turns pink. We lost a little bit of the KHP for the 4th trial. Trial 4 got a very very faint pink color. Even though we lost

some KHP, it looks near perfect titration. Trial 4 looked clear until held against white paper. Trial 5 was a dark pink, though better than 1, 2, and 3. 4th trial was the best out of the 5.

Trial 4







Revision #1

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